

SINGLE REPLACEMENT REACTION WORKSHEET

PRACTICE REACTION

1. $\text{Ag} + \text{KNO}_3 \rightarrow \text{No Reaction}$ (KNO_3 ~~is Ag~~)
2. $\text{Zn} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Zn}(\text{NO}_3)_2$
3. $2\text{Al} + 3\text{H}_2\text{SO}_4 \rightarrow \text{Al}_2(\text{SO}_4)_3 + 3\text{H}_2$
4. $\text{Cl}_2 + 2\text{KI} \rightarrow 2\text{KCl} + \text{I}_2$
5. $2\text{Li} + 2\text{H}_2\text{O} \rightarrow 2\text{LiOH} + \text{H}_2$
6. $\text{Cu} + \text{FeSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu}$ (NO Reaction)
7. $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$
8. $\text{Fe} + \text{Pb}(\text{NO}_3)_2 \rightarrow \text{Fe}(\text{NO}_3)_2 + \text{Pb}$
9. $\text{Cu} + \text{H}_2\text{O} \rightarrow \text{No Reaction}$
10. $\text{Cu} + \text{Al}_2(\text{SO}_4)_3 \rightarrow \text{Al}_2(\text{SO}_4)_3 + \text{Cu}$ (NO Reaction)
11. $2\text{Al} + 3\text{Pb}(\text{NO}_3)_2 \rightarrow 2\text{Al}(\text{NO}_3)_3 + 3\text{Pb}$
12. $\text{Cl}_2 + 2\text{NaI} \rightarrow 2\text{NaCl} + \text{I}_2$
13. $\text{Fe} + 2\text{Ag}(\text{CH}_3\text{COO}) \rightarrow \text{Fe}(\text{CH}_3\text{COO})_2 + 2\text{Ag}$
14. $2\text{Al} + 3\text{CuCl}_2 \rightarrow 2\text{AlCl}_3 + 3\text{Cu}$
15. $\text{Br}_2 + \text{CaI}_2 \rightarrow \text{CaBr}_2 + \text{I}_2$
16. $2\text{Al} + 6\text{HCl} \rightarrow 2\text{AlCl}_3 + 3\text{H}_2$
17. $\text{Mg} + \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$
18. $\text{Zn} + \text{H}_2\text{SO}_4 \rightarrow \text{ZnSO}_4 + \text{H}_2$
19. $\text{Fe} + \text{CuSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu}$
20. $\text{Cl}_2 + \text{MgI}_2 \rightarrow \text{MgCl}_2 + \text{I}_2$